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4.2.2 How much energy is released when bonds  
form to produce 2 mol of  $H_2O$  and 1 mol of

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CO 2? (4) 4.2.3 Hence calculate the energy transferred ( $\Delta H$ ) for the combustion of 1 mol CH<sub>4</sub>. (2) 4.2.4 Draw a labelled energy profile for this reaction. Include the labels: ENTHALPY CHANGE ( $\Delta H$ ) ACTIVATION ENERGY (E<sub>a</sub>) (6)

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